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### **Chapter 22**

### **Oxidation Reduction**

### **Reactions**

Chem 2: Chapter 22

Oxidation Reduction

Reactions. STUDY.

PLAY. Oxidation. The

complete or partial loss

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**Oxidation** is the loss of electrons from a substance involved in a redox reaction. Ex: burning, bleaching, or rusting. **Reduction** is the complete or partial gain of electrons or loss of oxygen. Ex: iron ore to iron.

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Transcript Chapter 22 -  
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Oxidation - Reduction

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Rxns 1 Section 22.1

The Meaning of  
Oxidation and  
Reduction 2

**OBJECTIVES:** - Define the terms “oxidation” and “reduction” in terms of electron loss or gain. - State the characteristics of a RedOx Rxn, and recognize oxidizing and reducing agents.

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Reduction Reactions  
Nature of Oxidation  
and Reduction  
Questions 1. Explain in  
your own words the  
following oxidation  
processes: a. Gain of  
oxygen b. Loss of  
hydrogen c. Loss of  
electrons 2. Explain in  
your own words the  
following reduction  
processes: a. Loss of  
oxygen b. Gain of  
hydrogen c. Gain of  
electrons 3.



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Oxidation

**Chemistry Student  
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Answer Key Chapter  
22 ...**

Chapter 22: Oxidation  
Reduction Reactions •

Oxidation number: A  
(+) or (-) number  
assigned to an atom -  
tells which atom is  
more electronegative •

Rules for assigning  
oxidation numbers o  
ON of an ion in a  
compound is usually  
equal to its

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Reactions 22.1 Oxygen  
in Reactions Practice  
Questions Answer the  
questions at the  
following link: Answers  
Answers are found  
here: Review Questions  
1. Give a definition of  
oxidation in terms of  
oxygen. 2. Give a  
definition of reduction  
in terms of oxygen. 3.

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xidation-Reduction\_  
Reactions\_-Answer**

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Chapter 16 - Covalent Bonding; Chapter 17 - Water and Aqueous Systems; Chapter 18 - Solutions; Chapter 19 - Reaction Rates & Equilibrium; Chapter 20 - Acids and Bases; Chapter 21 - Neutralization; Chapter 22 - Oxidation-Reduction Reactions; Chapter 23 - Electrochemistry; Chapter 24 - The Chemistry of Metals and Nonmetals;

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Chapter 25 ...

Reduction

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Chapter 22 -**

## **Oxidation-Reduction Reactions**

Oxidation-reduction (redox) reactions involve the transfer of electrons from one atom to another. Oxidation numbers are used to keep track of electrons in atoms. There are rules for assigning oxidation numbers to atoms.

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Redox (Oxidation &

Reduction) and

Electrochemistry 3

Chapter 22-23

Assignment & Problem

Set •Read Chapters 22

and 23, except all

students skip “Using

Oxidation Number

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Change" pp 663-664,  
and Regents students  
can skip section 23.2  
"Half-Cells and Cell  
Potentials" pp685-691.  
• Lab 24: Oxidation-  
Reduction Reactions

## **Redox (Oxidation & Reduction) and Electrochemistry Chapter ...**

Chapter 20 Notes  
Oxidation-Reduction  
Reactions 20.1 The  
Meaning of Oxidation  
and Reduction What

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are Oxidation and  
Reduction? o Oxygen  
and Redox KEY = The  
substance gaining O is  
oxidized, while the  
substance losing O is  
reduced Oxidation-  
Reduction Reactions =  
Reactions with a  
substance being  
oxidized and another ...

## **Chapter 20 Notes Oxidation-Reduction Reactions**

In this video I'll teach  
you about several

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Oxidation and  
Reduction reactions,  
which include carbonyl  
reductions using  
sodium borohydride,  
lithium aluminum  
hydride, and DIBAL-H  
(diisobutyl aluminum ...

## **Chapter 20 - Oxidation and Reduction Reactions: Part 1 of 2**

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Reactions "Redox  
Reactions" is the  
eighth chapter in the



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NCERT class 11 chemistry textbook. This chapter is regarded by many as one of the most important chapters in the CBSE class 11 chemistry syllabus, owing to the fact that the entire field of electrochemistry deals with redox reactions.

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## pyhounds.pridesour ce Reduction

A half reaction is an equation that shows just the oxidation or just the reduction that takes place in a redox reaction In this method you write and balance the oxidation half reaction, then you write and balance the reduction half reaction

In the example:  $S(s) + HNO_3 \rightarrow SO_2(g) + NO(g) + H_2O(l)$



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Oxidation half reaction

NO<sub>3</sub> Reduction

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