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Holt Modern Chemistry Chapter 7 The chemical formula for an ionic compound represents one formula unit - the simplest ratio of the compound's positive ions (cations) and its negative ions (anions). Monatomic cations are identified simply by the element's name.

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Modern Chemistry - Holt Chapter 7. Flashcard maker : Lily Taylor. chemical formula. subscript of each element reveals number of atoms contained in a single molecule of a certain compound. monoatomic ions. ions formed from a single atom; remember to change nonmetals to -ide ending. oxyanions.

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Holt McDougal Modern Chemistry Chapter 7 Section 3-4 Vocabulary. STUDY. PLAY. Formula mass. The formula unit of a molecule or ion is the sum of the average atomic masses of all atoms represented in its formula. Percentage composition.

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CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds SECTION 1 SHORT ANSWER Answer the following questions in the space provided.
1. c In a Stock system name such as iron (III) sulfate, the Roman numeral tells us (a) how many atoms of Fe are in one formula unit. (b) how many sulfate ions can be attached to the iron atom.

Holt Modern Chemistry Chapter 7 Section 1 Review Answers

7H 5(NO 2) 3. 4 elements a. How many elements make up this compound? 6 oxygen atoms b. How many oxygen atoms are present in one molecule of C 7H 5(NO 2) 3? 21 atoms c. How many atoms in total are present in one molecule of C 7H 5(NO 2) 3? 4.2 1024 atoms d. How many atoms are present in a sample of 2.0 1023 molecules of C 7H 5(NO 2) 3? 4.

7 Chemical Formulas and Chemical Compounds

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Holt Modern Chemistry Chapter 7 Review Answer Key

_ Name Modern Chemistry _ Date Class • CHAPTER 7 HOMEWORK 7-3 (pp.211-213) VOCABULARY Write the prefix that corresponds to each number.

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Try not to use the book. Prefix Number Number Prefix 1 6 2 7 3 8 4 9 5 10 GRAPHIC ORGANIZER Complete the chart by supplying the missing information.

Modern Chemistry • CHAPTER 7

CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Write formulas for the following compounds: CuCO₃ a. copper(II) carbonate Na₂SO₃ b. sodium sulfite (NH₄)₃PO₄ c. ammonium phosphate SnS₂ d. tin(IV) sulfide HNO₂ e. nitrous acid 2.

7 Chemical Formulas and Chemical Compounds

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